Manhattan Project - Problem Set 4 - Solution (2) (a) Momentum conservation says Pn, i + 0 = Pn, f + Pp, f $I(a) \xrightarrow{4}_{2} \chi + \xrightarrow{9}_{4} Be \rightarrow \xrightarrow{13}_{6} C$ Solve for Pn, f = Pn, i - Pp, f $z_1 = z$ $z_2 = 4$ $A_1 = 4$ $A_2 = 9$ (b) $E_{N,i} + E_{P,i} = E_{N,f} + E_{P,f}$ (b) $E_{coulomb} = \frac{e^2}{4\pi \epsilon_0 a_0} \frac{\overline{z_1 \overline{z_2}}}{A_1^{V_3} + A_2^{V_3}}$ $mc^{2} + \frac{P_{n,i}}{z_{M}} + mc^{2} = mc^{2} + \frac{P_{n,i}}{z_{M}} + mc^{2} + \frac{P_{p,f}}{z_{M}}$ $\frac{e^2}{4\pi\epsilon_0 a_0} = 1.2 \text{ MeV}$ $\frac{the two}{formulas}$ $\frac{formulas}{needed}$ All the MC² terms cance?! Then multiply through by Zm in what is left and you just have Now just plug in $E_{coulomb} = 1.2 MeV \frac{2.4}{4^{1/3} + 9^{1/3}} = 2.6 MeV$ $P_{n_{j}i}^{z} = P_{n,f}^{2} + P_{P,f}^{2}$ (c) $P_{n,i}^{z} = (P_{n,i} - P_{P,f})^{2} + P_{P,f}^{2}$ (c) yes. (5.3 MeV > 2.6 MeV) $P_{n,i} = P_{n,i} - Z_{Pn,i} P_{p,f} + P_{p,f} + P_{p,f}$ $O = - Z P n_i P_{p,f} + Z P_{p,f}$ Pn, iPp, f = Pp, f => Pn, i = Pp, f (e) Economb = 1.2 2.92 4"3+238"3 = 28.4 MeV (d) We can divide out Pp, f, but it might have been zero. That corresponds to the $P_{p,f} = 0$ (f) No $\implies Pn, f = Pn, i$ That is the case where nothing happened -the nextron missed the proton. (e) Pn, i = Pp, f. . (f) All the energy and momentum is transfered.

3. Fission Products (a) I get 0.5 2-0.5 4(c) 10°.4 = 2.5 So the plot is felling us that the absorbation cross section of a thermal neutron hitting U-238 is Z.5 barns •-//\ (6) $10^{0.5} = 3.16$ (c) About 3% of the time you get A=95 (d) I get about 2.7 $(e) 10^{2.7} = 501$ (1) I get -2.2 So the plot is telling us that the fission cross section $(e)_{10}^{-2.2} = -0.006$ for thermal neutrons hitting U-235 is sol borns $(f) \frac{e \notin I}{e \# I + c \# 140} = \frac{501}{501 + 2.5 \# 140}$ (f) About 0.01% of the time you get A=116 4. Relative Cross Sections $=\frac{501}{501+350}=\frac{501}{851}$ (a) $\log_{10} \frac{1}{40} = -1.6$ = 0.59 The fissions can come from (6) -6-1.6=-7.6 U-235 despite them being only 1 in every 140 atoms. $\left(\log\left(\frac{1}{40}\times10^{-6}\right)\right)$ I got about 0.4